

In an analogous set of experiments [time 5 h, solution amounts 4 ml, catalyst  $\text{Pt}(\text{PPh}_3)_2\text{O}_2$  and  $\text{Ir}(\text{CO})(\text{PPh}_3)_2\text{Cl}$ ], gas-chromatographic analysis of products of hydroperoxide decomposition was also carried out.

Taking the initial solution as reference it was established that in absence of the catalyst there is only a slight thermal decomposition of the hydroperoxide (the final active oxygen is in fact about 70% of the initial one) with a sharp tendency to polymer formation instead of the usual monomeric products; in presence of the catalyst on the other hand, there is a sharp decrease of the final active oxygen [to about 40% with  $\text{Pt}(\text{PPh}_3)_2\text{O}_2$  and to about 10% with  $\text{Ir}(\text{CO})(\text{PPh}_3)_2\text{Cl}$ , always compared to the initial content] with a parallel increase in the amount of cyclohexene oxide, 2-cyclohexen-1-one, and 2-cyclohexen-1-ol.

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